Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_\_

Study Guide Periodic Table/Matter Test

Define each of the following terms.

1. Atom
2. Electron
3. Proton
4. Neutron
5. Molecule
6. Compound
7. Valence Electron
8. Valence Shells
9. Lewis Dot Diagram
10. Bohr Model
11. Symbol
12. Metals
13. Nonmetals
14. Metalloids
15. Nucleus
16. What makes an atom unique?
17. What is the atomic number based on?

Find the atomic mass of each of the following atoms:

1. Li
2. Ar
3. Ag
4. He
5. C
6. Tell what is important about elements in a specific Group/column on the PTE.
7. Name two elements in Group 6. How many valence electrons do they have?
8. Tell what is important about elements in a specific Period/Row on the PTE.
9. Name two elements in Period 3. How many electron shells do they have?
10. Which side of the PTE is the most reactive?
11. Relative to the “zigzag” line, where are the metals?
12. Where are the non metals?
13. Name the metalloids
14. Are the majority of elements on the PTE metals, nonmetals, or metalloids?

Name the chemical formula for the next 4 compounds.

1. Water?
2. Salt?
3. Carbon dioxide?
4. Sulfuric acid?
5. Find the molecular mass of one molecule of each of the above molecules from the previous questions 32-35.
6. Which group of elements does not bond easily? Explain why?
7. Classify each as a metal, nonmetal, or metalloid
8. K
9. S
10. N
11. C
12. Ar
13. Ag
14. O
15. Ca
16. Na

How many valence electrons each of the above atoms have?

1. K
2. S
3. N
4. C
5. Ar
6. Ne
7. He
8. Ca
9. O

Tell the charges for the ions in:

1. NaO2
2. K2S

Name the following compounds:

1. NaCl
2. KBr
3. CaCl2
4. CO2
5. H2O
6. H2SO4
7. HO